

Quantitative Analysis - Titrations

The key areas of study in this topic are:

- How to carry out calculations involving concentrations
- How to carry out an accurate titration
- How to prepare soluble salts

By the end of this topic I should be able to:

	Start	End
3.16 Explain why, if soluble salts are prepared from an acid and a soluble reactant: <ul style="list-style-type: none"> • a titration must be used • the acid and the soluble reactant are then mixed in the correct proportions • the solution remaining, after reaction, is only salt and water 		
3.18 Describe how to carry out an acid-alkali titration, using burette, pipette and a suitable indicator, to prepare a pure, dry salt		
5.8C Calculate the concentration of solutions in mol dm ⁻³ and convert concentration in g dm ⁻³ into mol dm ⁻³ and vice versa		
<i>5.9C Core Practical: Carry out an accurate acid-alkali titration, using burette, pipette and a suitable indicator</i>		
5.10C Carry out simple calculations using the results of titrations to calculate an unknown concentration of a solution or an unknown volume of solution required		

