

Periodic Table

The key areas of study in this topic are:

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By the end of this topic I should be able to:

	Start	End
1.13 Describe how Mendeleev arranged the elements, known at that time, in a periodic table by using properties of these elements and their compounds		
1.14 Describe how Mendeleev used his table to predict the existence and properties of some elements not then discovered		
1.15 Explain that Mendeleev thought he had arranged elements in order of increasing relative atomic mass but this was not always true because of the relative abundance of isotopes of some pairs of elements in the periodic table		
1.16 Explain the meaning of atomic number of an element in terms of position in the periodic table and number of protons in the nucleus		
1.17 Describe that in the periodic table <ul style="list-style-type: none"> • elements are arranged in order of increasing atomic number, in rows called periods • elements with similar properties are placed in the same vertical columns called groups 		
5.1C Recall that most metals are transition metals and that their typical properties include: <ul style="list-style-type: none"> • high melting point • high density • the formation of coloured compounds • catalytic activity of the metals and their compounds as exemplified by iron 		