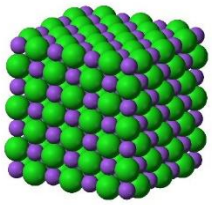
Enthalpy and Entropy



The key areas of study in this topic are:

- Entropy
- Free energy

By the end of this topic I should be able to:	Start	End
Explain that entropy is a measure of the dispersal of energy in a system which is greater, the more disordered a system		
Explain the difference in magnitude of the entropy of a system: • of solids, liquids and gases • for a reaction in which there is a change in the number of gaseous molecules		
Calculate the entropy change of a system, ΔS, and related quantities for a reaction given the entropies of the reactants and products		
Explain that the feasibility of a process depends upon the entropy change and temperature in the system, TΔS, and the enthalpy change of the system, ΔH		
State and use the relationship $\Delta G = \Delta H - T\Delta S$		
Explain that a process is feasible when ΔG has a negative value		
Understand the limitations of predictions made by ΔG about feasibility, in terms of kinetics		

In all topic areas you should be able to demonstrate and apply your knowledge and understanding.