

The Periodic Table: Group 2

The key areas of study in this topic are:

- Redox reactions and reactivity of Group 2 metals
- Reactions of Group 2 compounds

By the end of this topic I should be able to:

	Start	End
Understand that Group 2 elements have an outer shell electron configuration of s^2 and that these electrons are lost in redox reactions to form $2+$ ions		
Describe the relative reactivities of the Group 2 elements $Mg \rightarrow Ba$ shown by their redox reactions with: <ul style="list-style-type: none"> • Oxygen • Water • Dilute acids 		
Explain the trend in reactivity in terms of the first and second ionisation energies of Group 2 elements		
Describe the action of water on Group 2 oxides and the approximate pH of any resulting solutions, including the trend of increasing alkalinity		
Describe uses of some Group 2 compounds as bases, including equations, for example (but not limited to): <ul style="list-style-type: none"> • $Ca(OH)_2$ in agriculture to neutralise acid soils • $Mg(OH)_2$ and $CaCO_3$ as 'antacids' in treating indigestion. 		

In all topic areas you should be able to demonstrate and apply your knowledge and understanding.

